

**AMENDMENTS TO THE SPECIFICATION**

**Please amend the paragraph beginning at page 23, line 8 as follows:**

When the aqueous glycolic acid solution (A) has a calculated monomeric glycolic acid weight ratio of less than ~~0.06~~ 0.60, the yield of the final glycolic acid crystals becomes disadvantageously low. On the other hand, when the aqueous glycolic acid solution (A) has a calculated monomeric glycolic acid weight ratio of more than 1.00, problems arise not only in that the purity of the final glycolic acid crystals is lowered, but also in that the aqueous glycolic acid solution or the glycolic acid crystals-containing slurry which is obtained after the deposition of glycolic acid crystals from the aqueous glycolic acid solution becomes too viscous, such that the handling property thereof is lowered and it becomes difficult to separate the deposited glycolic acid crystals from the aqueous glycolic acid solution.

**Please amend the heading at page 59, line 1 as follows:**

BEST MODE FOR CARRYING OUT THE INVENTION

**Please amend the paragraph beginning at page 66, line 2 (to page 67, line 2) as follows:**

~~About 5~~ About 0.5 g of an aqueous glycolic acid solution (hereinafter, referred to as "feedstock liquid") was weighed and fed into a 50 ml-volumetric flask. Then, 0.1 g of n-dodecane was weighed and added to the aqueous glycolic acid solution as an internal standard. The resultant mixture in the volumetric flask was diluted with dehydrated pyridine to obtain 50 ml of a diluted solution. Then, 0.3 ml of the diluted solution was added to 1 ml of N,O-bis-trimethylsilylacetamide, and the resultant mixture was allowed to stand at room temperature for 1 hour, thereby obtaining a sample solution. The sample solution was analyzed by gas chromatography (GC) under the following conditions:

Column: DB-1 (trade name; manufactured and sold by J&W Scientific, U.S.A) (column length: 30 m, inner diameter: 0.25 mm, film thickness: 1  $\mu$ m);

Carrier gas: helium;

Detector: hydrogen flame ionization detector (FID);

Injection temperature: 250 °C;

Detector temperature: 300 °C; and

Column temperature: first, the column temperature was elevated from 50 to 100 °C at a rate of 10 °C/min and maintained at 100 °C for 10 minutes and, then, elevated to 250 °C at a rate of 10 °C/min and maintained at 250 °C for 15 minutes.

**Please amend the paragraph beginning at page 74, line 24 (to page 75, line 9) as follows:**

Separately from the above, a calibration curve was obtained as follows. With respect to the monodisperse standard polymethyl methacrylate (PMMA) samples having molecular weights of 1,577,000, 685,000, 333,000, 100,250, 62,600, 24,300, 12,700, 4,700 and 1680, respectively, the respective ~~elution~~ elution times were determined by refractive index (RI) detection. From the obtained values of the respective ~~elution~~ elution times, a calibration curve was obtained. Using the calibration curve, the weight average molecular weight of the polymer was determined from the ~~elution~~ elution time thereof.

**Please amend the paragraph beginning at page 85, line 15 (to page 86, line 3) as follows:**

Dried glycolic acid crystals were obtained in substantially the same manner as in Example 3, except that the deposited crystals were subjected to pressure filtration under 0.3 MPa (in which air was used as a pressurizing medium) by means of a pressure filtration apparatus using a filter paper (3250, manufactured and sold by Azumi Filter Paper Co., Ltd., Japan),

followed by dehydration of the resultant residue under a stream of air for 10 minutes (performed after completion of a continuous discharge of the filtrate), and ~~the that~~ that the recovered crystals were not washed. The pressure filtration of the slurry containing the deposited crystals and the separation of the crystals by dehydration under a stream of air were able to be performed efficiently.

**Please amend the paragraph beginning at page 95, line 16 as follows:**

On the other hand, the removal of impurities from the ~~crystal~~ crystals 2 was performed in the same manner as in the removal of impurities from the crystals 1 (i.e., by repeating a series of the above-mentioned operations (i.e., the washing, the pressure filtration and the dehydration under a stream of air) twice), and the resultant purified crystals 2 were dried in the same manner as in the drying of the purified crystals 1, thereby obtaining final dried glycolic acid crystals.

**Please amend the paragraph beginning at page 99, line 18 (to page 100, line 1) as follows:**

On the other hand, the removal of impurities from the ~~crystal~~ crystals 4 was performed in the same manner as in the removal of impurities from the crystals 3 (i.e., by repeating a series of the above-mentioned operations (i.e., the washing, the pressure filtration and the dehydration under a stream of air) twice), and the resultant purified crystals 4 were dried in the same manner as in the drying of the purified crystals 3, thereby obtaining final dried glycolic acid crystals.

**Please amend Table 1 at page 111 as follows:**

Table 1

Example	Purity of dried glycolic acid <del>crystal</del> <u>crystals</u> used (% by weight)	Solid-phase polymerization reaction time (hr)	Weight average molecular weight of glycolic acid-lactic acid copolymer resin produced
Resin Production Example 1	99.96	30	180,000
Resin Production Example 2	99.96	30	185,000
Resin Production Example 3	99.97	30	186,000
Resin Production Example 4	99.92	30	160,000
Resin Production Example 5	99.90	30	152,000
Resin Production Example 6	99.47	30	50,000
		50	49,000
Resin Production Example 7	99.55	30	55,000
		50	53,000
Resin Production Example 8	99.68	30	80,000
		50	81,000
Resin Production Example 9	99.69	30	81,000
		50	82,000
Resin Production Example 10	99.72	30	88,000
		50	88,000
Resin Production Example 11	99.70	30	84,000
		50	85,000
Resin Production Example 12	99.97	30	185,000
Resin Production Example 13	99.99	30	210,000
Resin Production Example 14	99.97	30	183,000
Resin Production Example 15	99.97	30	184,000